

### COVALENT – prefixes

(1)mono- (2)di- (3)tri- (4)tetra- (5)penta- (6)hexa- (7)hepta- (8)octa- (9)nona- (10)deca-

### COVALENT – diatomic elements

H<sub>2</sub> N<sub>2</sub> O<sub>2</sub> F<sub>2</sub> Cl<sub>2</sub> Br<sub>2</sub> I<sub>2</sub>

### IONIC – common ions and polyatomic ions

1+	2+	3+	4+
ammonium, NH <sub>4</sub> <sup>+</sup> cesium, Cs <sup>+</sup> copper(I), Cu <sup>+</sup> hydrogen, H <sup>+</sup> lithium, Li <sup>+</sup> potassium, K <sup>+</sup> rubidium, Rb <sup>+</sup> silver, Ag <sup>+</sup> sodium, Na <sup>+</sup>	barium, Ba <sup>2+</sup> beryllium, Be <sup>2+</sup> cadmium, Cd <sup>2+</sup> calcium, Ca <sup>2+</sup> chromium(II), Cr <sup>2+</sup> cobalt(II), Co <sup>2+</sup> copper(II), Cu <sup>2+</sup> iron(II), Fe <sup>2+</sup> lead(II), Pb <sup>2+</sup> magnesium, Mg <sup>2+</sup> manganese(II), Mn <sup>2+</sup> mercury(I) <i>or</i> dimercury, Hg <sub>2</sub> <sup>2+</sup> * mercury(II), Hg <sup>2+</sup> nickel(II), Ni <sup>2+</sup> strontium, Sr <sup>2+</sup> tin(II), Sn <sup>2+</sup> vanadium(II), V <sup>2+</sup> zinc, Zn <sup>2+</sup>	aluminum, Al <sup>3+</sup> antimony(III), Sb <sup>3+</sup> arsenic(III), As <sup>3+</sup> chromium(III), Cr <sup>3+</sup> cobalt(III), Co <sup>3+</sup> iron(III), Fe <sup>3+</sup> lead(III), Pb <sup>3+</sup> vanadium(III), V <sup>3+</sup>	lead(IV), Pb <sup>4+</sup> tin(IV), Sn <sup>4+</sup> vanadium(IV), V <sup>4+</sup>

1-	2-	3-	
acetate, C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> <i>or</i> CH <sub>3</sub> COO <sup>-</sup> bromate, BrO <sub>3</sub> <sup>-</sup> bromide, Br <sup>-</sup> chlorate, ClO <sub>3</sub> <sup>-</sup> chloride, Cl <sup>-</sup> chlorite, ClO <sub>2</sub> <sup>-</sup> cyanide, CN <sup>-</sup> dihydrogen phosphate, H <sub>2</sub> PO <sub>4</sub> <sup>-</sup> fluoride, F <sup>-</sup> hydride, H <sup>-</sup> hydrogen carbonate <i>or</i> bicarbonate, HCO <sub>3</sub> <sup>-</sup> hydrogen sulfate <i>or</i> bisulfate, HSO <sub>4</sub> <sup>-</sup> hydrogen sulfide, HS <sup>-</sup> hydroxide, OH <sup>-</sup> hypochlorite, ClO <sup>-</sup> iodate, IO <sub>3</sub> <sup>-</sup> iodide, I <sup>-</sup> nitrate, NO <sub>3</sub> <sup>-</sup> nitrite, NO <sub>2</sub> <sup>-</sup> perchlorate, ClO <sub>4</sub> <sup>-</sup> permanganate, MnO <sub>4</sub> <sup>-</sup> thiocyanate, SCN <sup>-</sup>	carbonate, CO <sub>3</sub> <sup>2-</sup> chromate, CrO <sub>4</sub> <sup>2-</sup> dichromate, Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup> hydrogen phosphate <i>or</i> monohydrogen phosphate, HPO <sub>4</sub> <sup>2-</sup> oxalate, C <sub>2</sub> O <sub>4</sub> <sup>2-</sup> oxide, O <sup>2-</sup> peroxide, O <sub>2</sub> <sup>2-</sup> sulfate, SO <sub>4</sub> <sup>2-</sup> sulfide, S <sup>2-</sup> sulfite, SO <sub>3</sub> <sup>2-</sup> tartrate, C <sub>4</sub> H <sub>4</sub> O <sub>6</sub> <sup>2-</sup>	arsenate, AsO <sub>4</sub> <sup>3-</sup> nitride, N <sup>3-</sup> phosphate, PO <sub>4</sub> <sup>3-</sup> phosphide, P <sup>3-</sup>	

### ACIDS - common

acetic acid, HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>(aq)

hydrochloric acid, HCl(aq)

nitric acid, HNO<sub>3</sub> (aq)

sulfuric acid, H<sub>2</sub>SO<sub>4</sub>(aq)